

BAB V

KESIMPULAN DAN SARAN

5.1 Kesimpulan

Dari hasil penelitian leaching *spent catalyst* Co-Mo/Al₂O₃, dapat disimpulkan:

1. Pada peningkatan waktu *leaching*, akan dihasilkan kecenderungan perolehan *recovery* ion logam yang meningkat pula akibat kontak antara pelarut dengan zat terlarut dalam kurun waktu yang lama dapat menghasilkan ekstrak yang semakin banyak.
2. Pada saat *column leaching*, nilai *recover* logam kobalt (Co²⁺) adalah yang terbesar (33,12%, 31,62%, dan 47,19%) yang diakibatkan oleh kelarutan ion kobalt akan semakin meningkat jika berada pada suasana asam.
3. Hasil ion logam molibdenum dan alumunium tidak terlalu besar jika dibandingkan dengan ion kobalt yang disebabkan karena ion logam molibdenum (Mo⁶⁺) dalam bentuk MoO₃ memiliki nilai kelarutan yang menurun pada suasana asam. Begitu juga pada ion logam alumunium (Al³⁺), kekuatan aluminium dalam membentuk ligan bersama senyawa organik asam sitrat rendah.
4. Pada variasi diameter partikel nilai *recovery* ion logam yang paling besar berada pada ukuran -20+40 mesh.
5. Pada model *Shrinking Core* diketahui bahwa yang mempengaruhi laju *leaching* adalah difusi secara internal akibat ikatan kompleks antara ligan dan ion logam yang terlalu besar.

5.2 Saran

Dari hasil penelitian leaching *spent catalyst* Co-Mo/Al₂O₃, dapat disarankan:

1. Menambah variasi laju alir pada *column leaching* sehingga didapatkan laju maksimum yang dapat digunakan pada column yang telah dirancang untuk menghasilkan nilai *recovery* ion logam yang optimum.
2. Menambah waktu *leaching* menjadi lebih dari dua hari sehingga dapat terlihat hasil *recovery* ion logam secara keseluruhan.
3. Menambah variasi asam organik sehingga dapat dilihat jenis asam organik yang cocok untuk *recovery spent catalyst* Co-Mo/Al₂O₃ pada *column leaching*.

4. Melakukan analisis XRF untuk hasil *spent catalyst* setelah *leaching* agar dapat diketahui komponen terkandung serta komposisinya.
5. Melakukan pengecekan nilai *flowrate* pada *head* pompa peristaltic yang akan masuk ke dalam setiap column
6. Melakukan pengecekan terhadap distribusi pada setiap ukuran partikel pada *column leaching*.
7. Melakukan pengecekan terhadap model kinetika selain *shrinking core model* sehingga dapat mengetahui model yang cocok pada penelitian ini.

DAFTAR PUSTAKA

- Al-Abed, S. R., Jegadeesan, G., Purandare, J., & Allen, D. (2008). Leaching behavior of mineral processing waste: Comparison of batch and column investigations. *Journal of Hazardous Materials*, 153(3), 1088–1092. <https://doi.org/10.1016/j.jhazmat.2007.09.063>
- Arslanoğlu, H., & Yaraş, A. (2019). Recovery of precious metals from spent Mo–Co–Ni/Al₂O₃ catalyst in organic acid medium: Process optimization and kinetic studies. *Petroleum Science and Technology*, 37(19), 2081–2093. <https://doi.org/10.1080/10916466.2019.1618867>
- Agnieszka SZUBERT, Michał ŁUPIŃSKI, Zygmunt SADOWSKI * APPLICATION OF SHRINKING CORE MODEL TO BIOLEACHING OF BLACK SHALE PARTICLES. (n.d.).
- Astuti, W., Hirajima, T., Sasaki, K., & Okibe, N. (2015). Kinetics of nickel extraction from Indonesian saprolitic ore by citric acid leaching under atmospheric pressure. *Minerals and Metallurgical Processing*, 32(3), 176–185. <https://doi.org/10.1007/bf03402286>
- Barik, S. P., Park, K. H., Parhi, P. K., & Park, J. T. (2012). Direct leaching of molybdenum and cobalt from spent hydrodesulphurization catalyst with sulphuric acid. *Hydrometallurgy*, 111–112(1), 46–51. <https://doi.org/10.1016/j.hydromet.2011.10.001>
- Barker, A. v., & Pilbeam, D. J. (2015). *Handbook Of Plant Nutrition* (2nd ed.). CRC Press.
- Beltran-Porter, A., Cervilla, A., Caturla, F., & Jose Vila, M. (1983). Lactate Complexes of Molybdenum (VI). *Transition Met. Chem.*, 8, 324–328.
- Bennett, C. R., McBride, D., Cross, M., & Gebhardt, J. E. (2012). A comprehensive model for copper sulphide heap leaching: Part 1 Basic formulation and validation through column test simulation. *Hydrometallurgy*, 127–128, 150–161. <https://doi.org/10.1016/j.hydromet.2012.08.004>
- Brown, L. S., & Holme, T. A. (2011). *Chemistry for Engineering Students* (2nd ed.). Brooks/Cole.
- Dawood, M. M.. 2012. “Recovery of Molybdenum from Spent Catalysts by Leaching Process.” *Thesis*. Nahrain University. Baghdad. Iraq.
- Doukeh, R., Bombos, D., Bombos, M., Oprescu, E. E., Dumitrascu, G., Vasilievici, G., & Calin, C. (2021). Catalytic hydrotreating of bio-oil and evaluation of main noxious emissions of gaseous phase. *Scientific Reports*, 11(1). <https://doi.org/10.1038/s41598-021-85244-z>
- Dusengemungu, L., Kasali, G., Gwanama, C., & Mubemba, B. (2021). Overview of Fungal Bioleaching of Metals. In *Environmental Advances* (Vol. 5). Elsevier Ltd. <https://doi.org/10.1016/j.envadv.2021.100083>

- Faraji, F., Alizadeh, A., Rashchi, F., & Mostoufi, N. (2020). Kinetics of leaching: A review. In *Reviews in Chemical Engineering* (Vol. 38, Issue 2, pp. 113–148). De Gruyter Open Ltd. <https://doi.org/10.1515/revce-2019-0073>,
- Febrian, M. B., Mulyati, T. S., Suherman, A., Adventini, N., Setiadi, Y., Setiawan, D., & Aziz, A. (2018). SPECTROPHOTOMETRIC DETERMINATION OF MOLYBDENUM CONTENT IN 99m Tc SOLUTION VIA Mo-TGA-KSCN COMPLEXES FORMATION. *Jurnal Sains Dan Teknologi Nuklir Indonesia*, 19(2), 71. <https://doi.org/10.17146/jstni.2018.19.2.4109>,
- Ferrier, R. J., Cai, L., Lin, Q., Gorman, G. J., & Neethling, S. J. (2015). Models for apparent reaction kinetics in heap leaching: a new semi-empirical approach and its comparison to shrinking core and other particle-scale models.
- Ghorbani, Y., Franzidis, J. P., & Petersen, J. (2015). Heap leaching technology - Current State, innovations, and future directions: A review. *Mineral Processing and Extractive Metallurgy Review*, 37(2), 73–119. <https://doi.org/10.1080/08827508.2015.1115990>,
- Golmohammadzadeh, R., Faraji, F., & Rashchi, F. (2018). Recovery of lithium and cobalt from spent lithium ion batteries (LIBs) using organic acids as leaching reagents: A review. In *Resources, Conservation and Recycling* (Vol. 136, pp. 418–435). Elsevier B.V. <https://doi.org/10.1016/j.resconrec.2018.04.024>
- Havlik, Tomas. 2008. *Hydrometallurgy Principles and Applications*. Cambridge.
- Haynes, W. M., Lide, D. R., & Bruno, T. J. (2017). *CRC Handbook of Chemistry and Physics* (97th ed.). CRC Press.
- He, Z., Zhang, Z., Yu, J., Zhou, F., Xu, Y., Xu, Z., Chen, Z., & Chi, R. (2016). Kinetics of column leaching of rare earth and aluminum from weathered crust elution-deposited rare earth ore with ammonium salt solutions. *Hydrometallurgy*, 163, 33–39. <https://doi.org/10.1016/j.hydromet.2016.02.016>
- Huang, J., Chen, M., Chen, H., Chen, S., & Sun, Q. (2014). Leaching behavior of copper from waste printed circuit boards with Brønsted acidic ionic liquid. *Waste Management*, 34(2), 483–488. <https://doi.org/10.1016/j.wasman.2013.10.027>
- Ilhan, S. (2017). Leaching of spent Ni–Mo hydrodesulphurization (HDS) catalyst in oxalic acid solutions. In *Minerals, Metals and Materials Series* (Issue 9783319521312, pp. 557–564). Springer International Publishing. https://doi.org/10.1007/978-3-319-52132-9_56,
- Ilhan, S. (2020). Extraction of Molybdenum, Nickel and Aluminium From Spent Ni–Mo Hydrodesulphurization (HDS) Catalyst in Oxalic Acid Solutions. *Canadian Metallurgical Quarterly*, 59(1), 26–35. <https://doi.org/10.1080/00084433.2020.1715691>
- Jackson, D. R., Garrett, B. C., & Bishop, T. A. (1984). *Comparison of Batch and Column Methods for Assessing Leachability of Hazardous Waste* (Vol. 18). <https://doi.org/10.1021/es00127a007>

- Jegan Roy, J., Srinivasan, M., & Cao, B. (2021). Bioleaching as an Eco-Friendly Approach for Metal Recovery from Spent NMC-Based Lithium-Ion Batteries at a High Pulp Density. *ACS Sustainable Chemistry and Engineering*, 9(8), 3060–3069. <https://doi.org/10.1021/acssuschemeng.0c06573>
- Kamalia, F. (2018). *SINTESIS KATALIS NiMo/γ-Al₂O₃, CoMo/γ-Al₂O₃, DAN CoNiMo/γ-Al₂O₃ UNTUK HIDRODESULFURISASI KEROSIN SKRIPSI FARAH KAMALIA PROGRAM STUDI KIMIA ALIA-FST.pdf*. <https://repository.uinjkt.ac.id/dspace/bitstream/123456789/47848/1/FARAH%20KAMALIA-FST.pdf>
- Kinraide, T. B., & Yermiyahu, U. (2007). A scale of metal ion binding strengths correlating with ionic charge, Pauling electronegativity, toxicity, and other physiological effects. *Journal of Inorganic Biochemistry*, 101(9 SPEC. ISS.), 1201–1213. <https://doi.org/10.1016/j.jinorgbio.2007.06.003>
- Kok Yong, H. (2010). *Electro-Bioleaching of Spent Hydroprocessing Ni-Mo Catalysts*. National University of Singapore.
- Komnitsas, K., Petrakis, E., Bartzas, G., & Karmali, V. (2019). Column leaching of low-grade saprolitic laterites and valorization of leaching residues. *Science of the Total Environment*, 665, 347–357. <https://doi.org/10.1016/j.scitotenv.2019.01.381>
- Le, M. N., & Lee, M. S. (2019). Leaching of Rare Metals from Spent Petroleum Catalysts by Organic Acid Solution. *Journal of the Korean Institute of Resources Recycling*, 28(6), 36–45. <https://doi.org/10.7844/kirr.2019.28.6.36>
- Li, L., Fan, E., Guan, Y., Zhang, X., Xue, Q., Wei, L., Wu, F., & Chen, R. (2017). Sustainable Recovery of Cathode Materials from Spent Lithium-Ion Batteries Using Lactic Acid Leaching System. *ACS Sustainable Chemistry and Engineering*, 5(6), 5224–5233. <https://doi.org/10.1021/acssuschemeng.7b00571>
- Li, L., Ge, J., Chen, R., Wu, F., Chen, S., & Zhang, X. (2010). Environmental friendly leaching reagent for cobalt and lithium recovery from spent lithium-ion batteries. *Waste Management*, 30(12), 2615–2621. <https://doi.org/10.1016/j.wasman.2010.08.008>
- Lison, D. (2015). Cobalt. In *Handbook on the Toxicology of Metals: Fourth Edition* (Vol. 1, pp. 743–763). Elsevier Inc. <https://doi.org/10.1016/B978-0-444-59453-2.00034-2>
- Liu, J., Qiu, Z., Yang, J., Cao, L., & Zhang, W. (2016). Recovery of Mo and Ni from Spent Acrylonitrile Catalysts Using An Oxidation Leaching-Chemical Precipitation Technique. *Hydrometallurgy*, 164, 64–70. <https://doi.org/10.1016/j.hydromet.2016.05.003>
- Liu, X., Cheng, J., Sprik, M., & Lu, X. (2013). Solution structures and acidity constants of molybdic acid. *Journal of Physical Chemistry Letters*, 4(17), 2926–2930. <https://doi.org/10.1021/jz401444m>

- Lin, C. L., Miller, J. D., & Garcia, C. (2005). Saturated flow characteristics in column leaching as described by LB simulation. *Minerals Engineering*, 18(10), 1045–1051. <https://doi.org/10.1016/j.mineng.2005.02.006>
- Lizama, H. M. (2004). A kinetic description of percolation bioleaching. *Minerals Engineering*, 17(1), 23–32. <https://doi.org/10.1016/j.mineng.2003.09.012>
- Marafi, M., & Stanislaus, A. (2008). Spent Catalyst Waste Management: A Review. Part I- Developments In Hydroprocessing Catalyst Waste Reduction and Use. In *Resources, Conservation and Recycling* (Vol. 52, Issue 6, pp. 859–873). <https://doi.org/10.1016/j.resconrec.2008.02.004>.
- McDonald, R. G., & Whittington, B. I. (2008). Atmospheric Acid Leaching of Nickel Laterites Review. Part I. Sulphuric Acid Technologies. *Hydrometallurgy*, 91(1–4), 35–55. <https://doi.org/10.1016/j.hydromet.2007.11.009>.
- Mgbeahuruike, L. U., Barrett, J., Potgieter, H. J., van Dyk, L., & Potgieter-Vermaak, S. S. (2019). A Comparison of Batch, Column and Heap Leaching Efficiencies for the Recovery of Heavy Metals from Artificially Contaminated Simulated Soil. *Journal of Environmental Protection*, 10(05), 632–650. <https://doi.org/10.4236/jep.2019.105038>
- Mitchell, K., Trakal, L., Sillerova, H., Avelar-González, F. J., Guerrero-Barrera, A. L., Hough, R., & Beesley, L. (2018). Mobility of As, Cr and Cu in a contaminated grassland soil in response to diverse organic amendments; a sequential column leaching experiment. *Applied Geochemistry*, 88, 95–102. <https://doi.org/10.1016/j.apgeochem.2017.05.020>
- Mohammed, A. (2016). *Synthesis Of Nano Ni-Mo/γ-Al₂O₃ Catalyst Viscosity index improvers for Lubricating Oil View project*
- Mohapatra, D., & Park, K. H. (2007). Selective Recovery of Mo, Co and Al from Spent Co/Mo/Γ -Al₂O₃ Catalyst: Effect of Calcination Temperature. *Journal of Environmental Science and Health - Part A Toxic/Hazardous Substances and Environmental Engineering*, 42(4), 507–515. <https://doi.org/10.1080/10934520601188409>.
- Mousavi, S. M., Yaghmaei, S., Vossoughi, M., Jafari, A., & Hoseini, S. A. (2005). Comparison of Bioleaching Ability of Two Native Mesophilic and Thermophilic Bacteria on Copper Recovery from Chalcopyrite Concentrate In An Airlift Bioreactor. *Hydrometallurgy*, 80(1–2), 139–144. <https://doi.org/10.1016/j.hydromet.2005.08.001>
- Muhammad, H., Wei, T., Cao, G., Yu, S. H., Ren, X. H., Jia, H. L., Saleem, A., Hua, L., Guo, J. K., & Li, Y. (2021). Study of soil microorganisms modified wheat straw and biochar for reducing cadmium leaching potential and bioavailability. *Chemosphere*, 273. <https://doi.org/10.1016/j.chemosphere.2021.129644>
- Nugroho, M. E. C., Sutijan, Prasetya, A., & Astuti, W. (2021). Recovery of Cobalt and Molybdenum from Spent Catalyst Using Citric Acid. *IOP Conference Series: Earth and Environmental Science*, 882(1). <https://doi.org/10.1088/1755-1315/882/1/012006>.

- Park, K. H., Mohapatra, D., Hong-In, K., & Xueyi, G. (2007). Dissolution behavior of a complex Cu-Ni-Co-Fe matte in CuCl₂-NaCl-HCl leaching medium. *Separation and Purification Technology*, 56(3), 303–310. <https://doi.org/10.1016/j.seppur.2007.02.013>,
- Patnaik, Pradyot. (2003). *Handbook Of Inorganic Chemicals*. McGraw-Hill.
- Pathak, A., Vinoba, M., & Kothari, R. (2021). Emerging Role of Organic Acids In Leaching of Valuable Metals from Refinery-Spent Hydroprocessing Catalysts, and Potential Techno-Economic Challenges: A Review. *Critical Reviews in Environmental Science and Technology*, 51(1), 1–43. <https://doi.org/10.1080/10643389.2019.1709399>,
- Petrucci, R. H., Geoffrey Herring, F., Madura, J. D., Bissonnette, C., & Canada Toronto, P. (2011). *General Chemistry Principles and Modern Applications* (D. Hunter, Ed.; 10th ed.). Pearson Education, Inc.
- Pornpetpaiboon, S., Lobthaisong, A., & Muncharoen, S. (2016). *Development of A Cobalt (II) Analysis Method In Laboratory Wastewaters Using Uv-Vis Spectrophotometry*. www.buuconference.buu.ac.th,
- Pradhan, D., Kim, D. J., Sukla, L. B., Pattanaik, A., & Lee, S. W. (2020). Evaluation of Molybdenum Recovery from Sulfur Removed Spent Catalyst Using Leaching and Solvent Extraction. *Scientific Reports*, 10(1). <https://doi.org/10.1038/s41598-020-58972-x>
- Pyper, R., Associates, C. &, Seal, T., Uhrie, J. L., & Miller, G. C. (2020). *Dump and Heap Leaching*.32.
- Quijada-Maldonado, E., Torres, M. J., & Romero, J. (2017). Solvent Extraction of Molybdenum (VI) from Aqueous Solution Using Ionic Liquids As Diluents. *Separation and Purification Technology*, 177, 200–206. <https://doi.org/10.1016/j.seppur.2016.12.045>
- Qiu, G., Li, Q., Yu, R., Sun, Z., Liu, Y., Chen, M., Yin, H., Zhang, Y., Liang, Y., Xu, L., Sun, L., & Liu, X. (2011). Column Bioleaching of Uranium Embedded In Granite Porphyry By A Mesophilic Acidophilic Consortium. *Bioresource Technology*, 102(7), 4697–4702. <https://doi.org/10.1016/j.biortech.2011.01.038>,
- Rahman Kustiawan, U., & Pratiwi, R. (2016). *Farmaka Dithizon: Agen Pengkompleks Untuk Analisis Logam Menggunakan Spektrofotometri Uv-Vis*. <https://doi.org/https://doi.org/10.24198/jf.v14i2.10899>,
- Reda, M. R. (1991). Regeneration of Spent Hydroprocessing Catalysts. 1. Effect of the Iron(II)/Iron(III) Redox Couple on the Selectivity of the Removal of Metals. In *Ind. Eng. Chem. Res* (Vol. 30).
- Rezki, A. S., Sumardi, S., Astuti, W., Made Bendiyasa, I., & Petrus, H. T. B. M. (2021). Molybdenum Extraction from Spent Catalyst Using Citric Acid: Characteristic and

- Kinetics Study. *IOP Conference Series: Earth and Environmental Science*, 830(1), 1–10. <https://doi.org/10.1088/1755-1315/830/1/012020>
- Rivas, E., Martín-Lara, M. Á., Blázquez, G., Pérez, A., & Calero, M. (2019). Column leaching tests to valorize a solidwaste from the decommissioning of coal-fired power plants. *Energies*, 12(9). <https://doi.org/10.3390/en12091684>
- Rouessac, F., Rouessac, A., & Brooks, S. (2007). *Chemical Analysis Modern Instrumentation Methods and Techniques Second Edition* (7th ed.). John Wiley & Sons Ltd.
- Saldaña, M., Gálvez, E., Robles, P., Castillo, J., & Toro, N. (2022). Copper Mineral Leaching Mathematical Models—A Review. In *Materials* (Vol. 15, Issue 5). MDPI. <https://doi.org/10.3390/ma15051757>.
- Setiawan, H., Petrus, H. T. B. M., & Perdana, I. (2019). Reaction kinetics modeling for lithium and cobalt recovery from spent lithium-ion batteries using acetic acid. *International Journal of Minerals, Metallurgy and Materials*, 26(1), 98–107. <https://doi.org/10.1007/s12613-019-1713-0>
- Sheta, M. E. (2008). *A new Leaching System, Sheta Extractor*.
- Shokrollahi, A., Ghaedi, M., Niband, M. S., & Rajabi, H. R. (2008). Selective and Sensitive Spectrophotometric Method For Determination of Sub-Micro-Molar Amounts of Aluminium Ion. *Journal of Hazardous Materials*, 151(2–3), 642–648. <https://doi.org/10.1016/j.jhazmat.2007.06.037>
- Skoog. 2014. “Fundamentals of Analytical Chemistry-Cengage Lea.”
- Solomons, T. W. G., Craig B. Fryhle, & Scott A. Snyder. (2008). *Solomons Organic Chemistry 11th Ed.*
- Slack, J. F., Kimball, B. E., & Shedd, K. B. (2017). Cobalt, chap. F of Schulz, K.J., DeYoung, J.H., Jr., Seal, R.R., II, and Bradley, D.C., eds., Critical mineral resources of the United States—Economic and environmental geology and prospects for future supply. *U.S. Geological Survey Professional Paper 1802*, F1–F40. <https://doi.org/10.3133/pp1802F>
- Srichandan, H., Mohapatra, R. K., Singh, P. K., Mishra, S., Parhi, P. K., & Naik, K. (2020). Column bioleaching applications, process development, mechanism, parametric effect and modelling: A review. In *Journal of Industrial and Engineering Chemistry* (Vol. 90, pp. 1–16). Korean Society of Industrial Engineering Chemistry. <https://doi.org/10.1016/j.jiec.2020.07.012>
- Szymczycha-Madeja, A. (2011). Kinetics of Mo, Ni, V and Al leaching from a spent hydrodesulphurization catalyst in a solution containing oxalic acid and hydrogen peroxide. *Journal of Hazardous Materials*, 186(2–3), 2157–2161. <https://doi.org/10.1016/j.jhazmat.2010.11.120>

- Wang, X., Sun, T., Chen, C., & Hu, T. (2017). Current Studies of Treating Processes for Nickel Laterite Ores. *Advances in Computer Science Research*, 70, 139–152.
- Wanta, K. C., Widi Astuti, Himawan Tri Bayu Murti Petrus, dan Indra Perdana. (2022). Product Diffusion-Controlled Leaching of Nickel Laterite using Low Concentration Citric Acid Leachant at Atmospheric Condition. *International Journal of Technology (IJTech)*. <https://doi.org/10.14716/ijtech.v13i2.4641>
- Wanta, K. C., Natapraja, E. Y., Susanti, R. F., Gemilar, G. P., Astuti, W., & Petrus, H. T. B. M. (2021). INCREASING OF METAL RECOVERY IN LEACHING PROCESS OF SPENT CATALYST AT LOW TEMPERATURE: THE ADDITION OF HYDROGEN PEROXIDE AND SODIUM CHLORIDE. *Metalurgi*, 36(2). <https://doi.org/10.14203/metalurgi.v36i2.591>,
- Wanta, K., B M Petrus, H. T., Perdana, I., Widi Astuti, dan, Penelitian Teknologi Material, B., Ilmu Pengetahuan Indonesia Jalan Ir Sutami Km, L., Bintang, T., & Selatan, L. (2017). Uji Validitas Model Shrinking Core terhadap Pengaruh Konsentrasi Asam Sitrat dalam Proses Leaching Nikel Laterit. 11(1), 30–35.
- Wanta, K. C., Perdana, I., & Petrus, H. T. B. M. (2016). Evaluation of shrinking core model in leaching process of Pomalaa nickel laterite using citric acid as leachant at atmospheric conditions. IOP Conference Series: Materials Science and Engineering, 162(1). <https://doi.org/10.1088/1757-899X/162/1/012018>
- Zinoveev, D., Pasechnik, L., Fedotov, M., Dyubanov, V., Grudinsky, P., & Alpatov, A. (2021). Extraction of valuable elements from red mud with a focus on using liquid media—a review. In *Recycling* (Vol. 6, Issue 2). MDPI AG. <https://doi.org/10.3390/recycling6020038>
- Zumdahl, S. S., & Zumdahl, S. A. (2014). *Chemistry* (9th ed.). Brooks Cole.